

# Moles

1. First Name
2. Last Name
3. Noun
4. Noun
5. Number
6. Noun
7. Number
8. Noun
9. Noun
10. Same Noun As Two Ago
11. Same Noun As Last
12. Number
13. Same Number As Last
14. Number
15. Number
16. Noun
17. Noun

# Moles

\_\_\_\_\_ Last name was an Italian \_\_\_\_\_ Noun who provided us with his number  $6.02 \times 10^{23}$ . This number shows us how many molecules in a mole. Every element has a molar mass, and in turn every compound has a molar mass. For example \_\_\_\_\_ Noun molar mass is \_\_\_\_\_ Number g/mol but \_\_\_\_\_ Noun molar mass is \_\_\_\_\_ Number g/mol. The molar mass is the same as the atomic number of the element. You can use the molar mass to find \_\_\_\_\_ Noun in \_\_\_\_\_ Noun and vice-versa. The way to do \_\_\_\_\_ Same noun as two ago to \_\_\_\_\_ Same noun as last is by multiplying the molar mass by the amount of moles and to find moles in mass you divide the atoms over molar mass. Molar volume is equal to 22.4L of one mole of an ideal gas at STP. STP is Standard Temperature and (which is 0° C or 32 ° F) and Pressure (which is 1 atm, sea level)> Like molar mass, molar volume can be used to find volume in moles and vice-versa. An example would be having \_\_\_\_\_ Number of O<sub>2</sub> at STP and you wanted to know how many moles that is. You would divide \_\_\_\_\_ Same number as last by \_\_\_\_\_ Number to get \_\_\_\_\_ Number moles. To get moles into volume you would multiply the amount of moles by 22.4L. Lastly, there is percentage composition. This is when you have a \_\_\_\_\_ Noun and you take the weight of one \_\_\_\_\_ Noun and divide it by the other, then you put the decimal into percentage.