

Moles

1. First Name
2. Last Name
3. Noun
4. Noun
5. Number
6. Noun
7. Number
8. Noun
9. Noun
10. Same Noun As Two Ago
11. Same Noun As Last
12. Number
13. Same Number As Last
14. Number
15. Number
16. Noun
17. Noun

Moles

_____ Last name _____ was an Italian _____ Noun _____ who provided us with his number 6.02×10^{23} . This number shows us how many molecules in a mole. Every element has a molar mass, and in turn every compound has a molar mass. For example _____ Noun _____ molar mass is _____ Number _____ g/mol but _____ Noun _____ molar mass is _____ Number _____ g/mol. The molar mass is the same as the atomic number of the element. You can use the molar mass to find _____ Noun _____ in _____ Noun _____ and vice-versa. The way to do _____ Same noun as two ago _____ to _____ Same noun as last _____ is by multiplying the molar mass by the amount of moles and to find moles in mass you divide the atoms over molar mass. Molar volume is equal to 22.4L of one mole of an ideal gas at STP. STP is Standard Temperature and (which is 0? C or 32 ? F) and Pressure (which is 1 atm, sea level)> Like molar mass, molar volume can be used to find volume in moles and vice-versa. An example would be having _____ Number _____ of O₂ at STP and you wanted to know how many moles that is. You would divide _____ Same number as last _____ by _____ Number _____ to get _____ Number _____ moles. To get moles into volume you would multiply the amount of moles by 22.4L. Lastly, there is percentage composition. This is when you have a _____ Noun _____ and you take the weight of one _____ Noun _____ and divide it by the other, then you put the decimal into percentage.