

# Light and Electrons

1. First Name Of A Person
2. Adverb
3. Adjective
4. Verb - Base Form
5. Adjective
6. Noun
7. Verb - Present Ends In Ing
8. Adjective
9. Adjective
10. Noun
11. Verb - Present Ends In S
12. Adjective
13. Noun - Plural

# Light and Electrons

The evidence used to support \_\_\_\_\_ model came from the atomic spectra. He suggested that an atomic spectrum is made by the electrons in an atom moving energy levels. The electrons \_\_\_\_\_ have the \_\_\_\_\_ energy possible, called ground state. If the electrons are given energy (through heat, electricity, light, etc) the electrons in an atom could \_\_\_\_\_ energy by jumping to a higher energy level or an \_\_\_\_\_ state. The electrons then give off the energy they had absorbed in the form of a piece of light, called a \_\_\_\_\_, to fall back to a lower energy level.

The energy emitted by electrons \_\_\_\_\_ back to \_\_\_\_\_ energy levels would always be precise amounts of energy because the differences in energy levels were precise. This explains why you see \_\_\_\_\_ lines of light when looking at an atomic spectrum - each line of light matches a specific "step down" that an \_\_\_\_\_ can take in that atom. This also \_\_\_\_\_ why each element produces a \_\_\_\_\_ atomic spectrum. Because each element has different acceptable energy levels for their electrons, the possible steps each element's electrons can take differ from all other \_\_\_\_\_.