Light and Electrons

1. Adjective _____
2. Noun - Plural _____
3. Adverb _____
4. Adjective _____
5. Adjective _____
6. Noun _____
7. Adjective _____
8. Verb - Present Ends In Ing _____
9. Verb - Present Ends In S _____
10. Adjective _____
11. Noun _____
12. Adjective _____
13. Adjective _____

Light and Electrons

The evidence used to support Bohr's model came from the _____Adjective_____ spectra. He suggested that an atomic spectrum is made by the electrons in an atom moving energy _____Noun - Plural_____. The electrons _____Adverb_____ have the _____Adjective_____ energy possible, called ground state. If the electrons are given energy (through heat, electricity, light, etc) the electrons in an atom could absorb energy by jumping to a higher energy level or _____Adjective_____ state. The electrons then give off the energy they had absorbed in the form of a piece of light, called a _____Noun_____, to fall back to a _____Adjective_____ energy level.

The energy emitted by electrons _____Verb - Present ends in ING_____ back to lower energy levels would always be precise amounts of energy because the differences in energy levels were precise. This _____Verb - Present ends in S_____ why you see _____Adjective_____ lines of _____Noun_____ when looking at an atomic spectrum - each line of light matches a specific "step down" that an electron can take in that atom. This also explains why each element produces a _____Adjective_____ atomic spectrum. Because each element has different acceptable energy levels for their electrons, the _____Adjective_____ steps each element's electrons can take differ from all other elements.